

INTRODUCTION

Chemistry is the study of matter & change

Mixtures - combination of substances combined physically, not chemically

Scientific notation

1. Write the number in scientific notation
2. Move the decimal point
3. Count the number of places moved

DIMENSIONAL ANALYSIS

BASE - substance that reacts with acids to form water



Acid

Substance that reacts with bases to form water

Stoichiometry

BALANCING CHEMICAL EQUATIONS

General Chemistry 1 Notes

kinetic theory GASES

30 PAGES

PERCENT YIELD

How successful the reaction was

$\% \text{ yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$

Acid-base reaction

Acids normally start with H or HO

ACIDS vs BASES

| ACIDS | BASES |
|--|---|
| <ul style="list-style-type: none"> release H^+ conduct electricity | <ul style="list-style-type: none"> release OH^- conduct electricity |

MOLECULAR EQUATIONS, IONIC EQUATIONS, NET IONIC EQUATIONS

Molecular Equations
Formulas written as if an species existed as molecules
Ex: $\text{Fe(NO}_3)_3(aq) + 2\text{I}^-(aq) \rightarrow \text{FeI}_3(s) + 2\text{NO}_3^-(aq)$

Identify the spectator ions

steps for writing ionic & net ionic equations

1. Write balanced molecular equation for the reaction, correct formulas & solubility

thermochemistry

the study of heat change

STATE FUNCTIONS: internal energy, pressure, temperature, volume, enthalpy

ENERGY, WORK & HEAT

OXIDATION-REDUCTION

OXIDATION: loss of electrons

DALTON'S LAW OF PARTIAL PRESSURES
Total pressure = sum of partial pressures

BALANCING STANDARD ENTHALPY OF FORMATION

HESS'S LAW: when reactants are converted into product, change in enthalpy is the same whether the reaction occurs in one or multiple steps

$$\Delta H_{\text{rxn}}^\circ = \sum n \Delta H_f^\circ(\text{products}) - \sum m \Delta H_f^\circ(\text{reactants})$$

1. multiply each coefficient by the number of molecules

2. $\text{H}_2\text{O}(l)$

= heat

$$\Delta H = [2 \Delta H_f^\circ(\text{CO}_2) + (2 \Delta H_f^\circ(\text{H}_2\text{O}))]$$

$$- [2 \Delta H_f^\circ(\text{C}_2\text{H}_6) + (7 \Delta H_f^\circ(\text{O}_2))]$$

SPECIFIC HEAT CAPACITY

heat capacity

Notes On General Chemistry

R Pring



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