

Polar covalence

refers to

unequal sharing of valence electrons between bonded atoms

extreme cases are

caused by

resulting in

occurs in

differing electronegativities of the atoms

a polar bond

any hetero-atomic molecule

Lewis acid-base adducts

ionic solids

which still illustrate the principle that bonding occurs when

electrons are simultaneously close to two nuclei

which can give rise to a

permanent dipole moment

not cancelled by other polar bonds

if

Polar Covalence

M Lipman



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