

A REGENCY ROMANCE

The
Nabob's
Daughter

JESS HEILEMAN

Nabobs Daughter

Lauren Gardner



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web limiting reactant theoretical and percentage yields key $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$ g if a reaction vessel contains 0 mol KO_2 and 0 mol H_2O what is the limiting reactant

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 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H₂S and multiply it with 2 mole H₂O over 2 mole H₂S from the equation
 $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} \times 18 \text{ g H}_2\text{O} = 149.4 \text{ g H}_2\text{O}$
 is your theoretical yield so percent yield = $\frac{137.1 \text{ g H}_2\text{O}}{149.4 \text{ g H}_2\text{O}} \times 100 = 91.77\%$
 problems 6 & 8 are similar to problem 5

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web limiting reagents problem one 1 for the balanced equation
 $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$
 if the reaction of 21.3 grams of O₂ produces a 47.3% yield how many grams of CO would be produced
 step one figure out how many moles of the limiting reagent you have
 $21.3 \text{ g O}_2 \times \frac{1 \text{ mole}}{32 \text{ g}} = 0.666 \text{ moles O}_2$
 $0.666 \text{ moles O}_2 \times \frac{2 \text{ moles CO}}{3 \text{ moles O}_2} = 0.444 \text{ moles CO}$
 $0.444 \text{ moles CO} \times 28 \text{ g/mole} = 12.4 \text{ g CO}$

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web percent yield is a measure of the efficiency of the experimental design
 yield efficiency = $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$
 in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction
 $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
 4 Al + 3 O₂ → 2 Al₂O₃
 b determine the theoretical yield of Al₂O₃
 $160.0 \text{ g Al} \times \frac{1 \text{ mole Al}}{27 \text{ g Al}} \times \frac{2 \text{ moles Al}_2\text{O}_3}{4 \text{ moles Al}} \times 101.96 \text{ g Al}_2\text{O}_3 = 152.8 \text{ g Al}_2\text{O}_3$
 $302.3 \text{ g Al}_2\text{O}_3 \times \frac{26.98 \text{ g Al}}{101.96 \text{ g Al}_2\text{O}_3} = 78.2 \text{ g Al}$
 c determine the percent yield
 $\frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

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web 7 Na₂C₂O₄ is the limiting reactant percent yield 86.68% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless
 10 nitrogen is the limiting reagent
 11 yes methane is the limiting reagent
 12 C is the limiting reagent
 4.33 g of H₂ are left over
 13

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remains when the reaction is complete

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web the percent yield is calculated as follows $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ label 3 7 3

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