

1. Draw the Lewis Structure for NO



2. Draw the Lewis Structure for NOF



3. What is the structural geometry of the NOF molecule
bent

4. The rate constant (k) for a reaction was measured as a function of a temperature. A plot of $\ln k$ versus $1/T$ (in K) is linear and has a slope of -7002 K . Calculate the activation energy for the reaction

$$k = e^{-E_a/R} (+)$$

↑ slope

$$\begin{aligned} -7002 &= -E_a/R \\ -7002 \cdot 8.314 &= -E_a \\ -58210 &= -E_a \\ E_a &= 5.821 \times 10^4 \text{ J/mol} \end{aligned}$$

5. The gas phase rxn of NO w/ F_2 to form NOF and F has an E_a of 6.3 kJ/mol and a frequency factor of $A = 6.00 \times 10^8 \text{ M}^{-1}\text{s}^{-1}$. The rxn is believed to be bimolecular:

$\text{NO}(g) + \text{F}_2(g) \rightarrow \text{NOF}(g) + \text{F}(g)$
What NO is a free radical. Free radicals, species with an odd number of unbound electrons, tends to be more reactive. This is one reason why NO reacts so readily with gases such as fluorine or chlorine.

What is the rate constant at 665°C ?

$$\begin{aligned} 665 + 273 &= 938 \text{ K} \\ k &= 6.00 \times 10^8 \cdot e^{(-6300 / 8.314 \cdot 938)} \\ k &= 2.67 \times 10^8 \end{aligned}$$

6. The rate constant of a chemical reaction increased from 0.100 s^{-1} to 2.80 s^{-1} upon raising the temperature from 25.0°C to 53.0°C .

a. Calculate the value of $(1/T_1 - 1/T_2)$ where T_1 is the initial and T_2 is the final.

$$\begin{aligned} 25 + 273 &= 298 \text{ K} \\ 53 + 273 &= 326 \text{ K} \\ 1/298 - 1/326 &= 2.88 \times 10^{-4} \text{ K}^{-1} \end{aligned}$$

Mastering Chemistry Problems

Michal Rosen-Zvi



Mastering Chemistry Problems:

Mastering Chemistry Problems Book Review: Unveiling the Power of Words

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Table of Contents Mastering Chemistry Problems

1. Understanding the eBook Mastering Chemistry Problems
 - The Rise of Digital Reading Mastering Chemistry Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Mastering Chemistry Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mastering Chemistry Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mastering Chemistry Problems
 - Personalized Recommendations
 - Mastering Chemistry Problems User Reviews and Ratings
 - Mastering Chemistry Problems and Bestseller Lists
5. Accessing Mastering Chemistry Problems Free and Paid eBooks

- Mastering Chemistry Problems Public Domain eBooks
 - Mastering Chemistry Problems eBook Subscription Services
 - Mastering Chemistry Problems Budget-Friendly Options
6. Navigating Mastering Chemistry Problems eBook Formats
- ePub, PDF, MOBI, and More
 - Mastering Chemistry Problems Compatibility with Devices
 - Mastering Chemistry Problems Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Mastering Chemistry Problems
 - Highlighting and Note-Taking Mastering Chemistry Problems
 - Interactive Elements Mastering Chemistry Problems
8. Staying Engaged with Mastering Chemistry Problems
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mastering Chemistry Problems
9. Balancing eBooks and Physical Books Mastering Chemistry Problems
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mastering Chemistry Problems
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mastering Chemistry Problems
- Setting Reading Goals Mastering Chemistry Problems
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mastering Chemistry Problems
- Fact-Checking eBook Content of Mastering Chemistry Problems
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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