

# HOW TO WRITE BLOG POSTS USING AI

KEEP WRITING EVEN IF YOU'RE SHORT ON TIME OR HAVE WRITER'S BLOCK



# Low Budget Way To Write Blog Posts Using AI Organically BATCH13 794

**SA Dillow**



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web mass 32 g mass 16 g mass 3 2 g 2 for jars d e f use the number of molecules in the circle to determine the number of

moles of  $O_2$  in the jar then calculate the mass of the gas in the jar  
 $d = 12 \text{ mol } O_2$   
 $e = 2 \text{ mol } O_2$   
 $f = 2 \text{ mol } O_2$   
 $g = 6 \text{ mol } O_2$   
 $h = 2 \text{ mass } 384 \text{ g}$   
 $i = \text{mass } 64 \text{ g}$   
 $mass 192 \text{ g}$   
 jars  $g$  and  $i$  contain ammonia gas  $NH_3$

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web solution average atomic mass of  $Li = 6.941 \text{ amu}$   $x = 6.0151 \text{ amu}$   $1 = 7.0160 \text{ amu}$   $6.941 = 1.0009x$   $7.0160 = 1.0009x + 0.075$   $x = 0.075$  corresponds to a natural abundance of  $6Li$  of  $7.5$  percent the natural abundance of  $7Li$  is  $x = 0.925$  or  $92.5$  percent  $3.7$  the unit factor required is  $6.022$

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web a molar mass  $Li_2CO_3 = 2 \times 6.941 \text{ g} + 12.01 \text{ g} + 3 \times 16.00 \text{ g} = 73.89 \text{ g}$  b molar mass  $CS_2 = 12.01 \text{ g} + 2 \times 32.07 \text{ g} = 76.15 \text{ g}$  c molar mass  $CHCl_3 = 12.01 \text{ g} + 1.008 \text{ g} + 3 \times 35.45 \text{ g} = 119.37 \text{ g}$  d molar mass  $C_6H_8O_6 = 6 \times 12.01 \text{ g} + 8 \times 1.008 \text{ g} + 6 \times 16.00 \text{ g} = 176.12 \text{ g}$  e molar mass  $KNO_3 = 39.10 \text{ g} + 14.01 \text{ g} + 3 \times 16.00 \text{ g} = 101.11 \text{ g}$

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