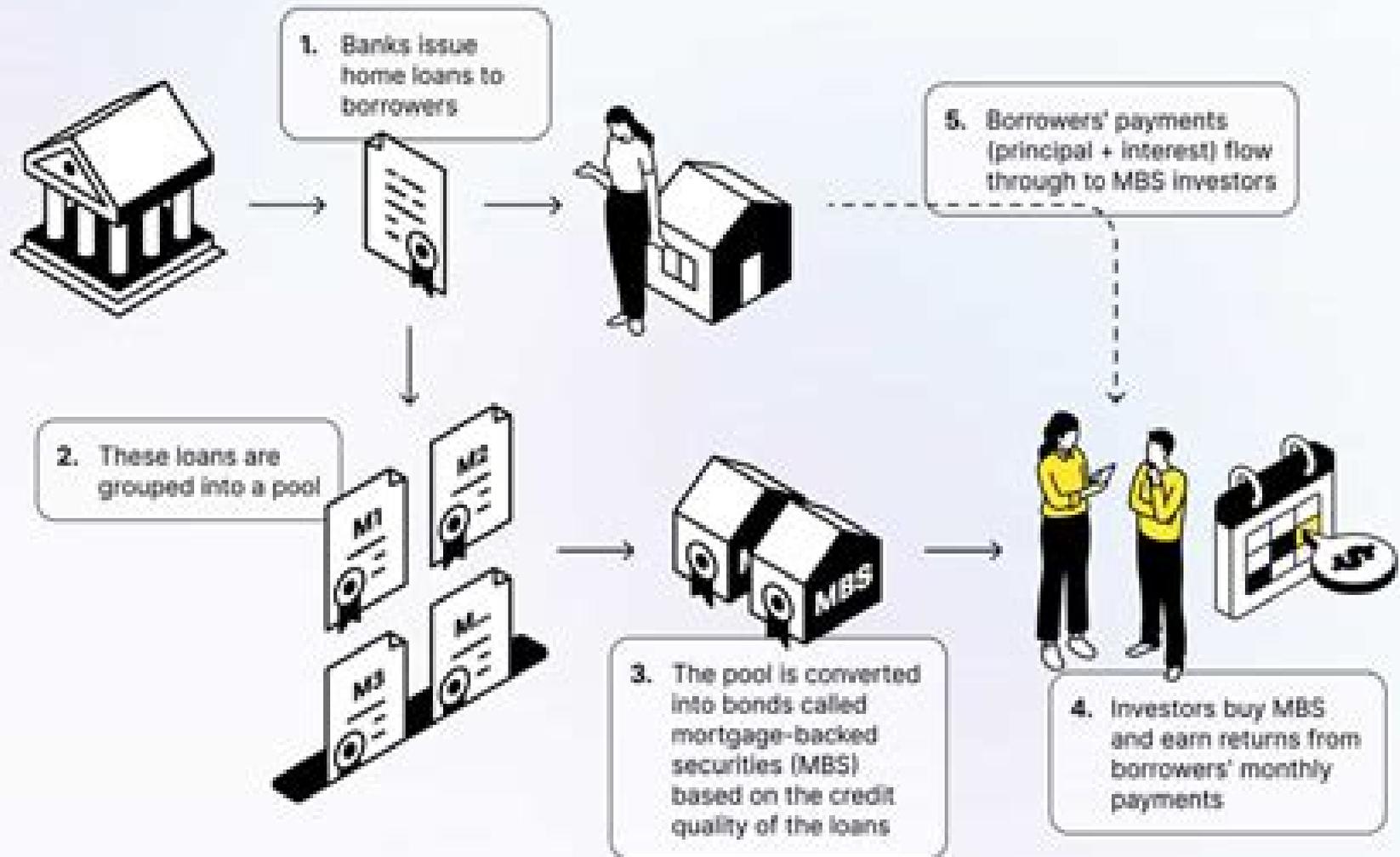


# Mortgage-Backed Securities for Dummies



# Mortgage Backed Securities

**D Siedentop**



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web limiting reactant theoretical and percentage yields key  $4 KO_2 + 2 H_2O \rightarrow 4 KOH + O_2$  g if a reaction vessel contains 0

mol ko 2 and 0 mol h 2 o what is the limiting reactant

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  

$$8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$$
  
 8.3 mole  $\text{H}_2\text{S}$  x 18g  $\text{H}_2\text{O}$  / 1 mole  $\text{H}_2\text{O}$  = 149.4g  $\text{H}_2\text{O}$  is your theoretical yield so percent yield =  $\frac{137.1\text{g } \text{H}_2\text{O}}{149.4\text{g } \text{H}_2\text{O}} \times 100 = 91.77\%$   
 problems 6 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts/html/limiting\_reagents\_theoretical\_actual\_and\_percent\_yields.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of  $\text{NH}_3$  can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation  

$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$
  
 if the reaction of 21.3 grams of  $\text{O}_2$  produces a 47.3% yield how many grams of CO would be produced step one figure out how many moles of the limiting reagent you have  

$$\frac{21.3 \text{ grams } \text{O}_2}{32 \text{ grams/mole}} = 0.666 \text{ moles } \text{O}_2$$
  

$$\frac{21.3 \text{ grams } \text{C}_3\text{H}_6}{42 \text{ grams/mole}} = 0.507 \text{ moles } \text{C}_3\text{H}_6$$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency =  $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  $\text{Al}_2\text{O}_3 + 4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160 g  $\text{Al}$  x 1 mol  $\text{Al}$  x 2 mol  $\text{Al}_2\text{O}_3$  x 101.964 g  $\text{Al}_2\text{O}_3$  302.3 g  $\text{Al}_2\text{O}_3$  26.982 g  $\text{Al}$  4 mol  $\text{Al}$  1 mol  $\text{Al}_2\text{O}_3$  c determine the percent yield  $\frac{\text{actual}}{\text{theoretical}} \times 100$   $\frac{260.0}{302.3} \times 100 = 86.01\%$

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h 2 are left over 13

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web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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80.0 g  $I_2O_5$  1 mol  $I_2O_5$  1 mol  $I_2$  x 1 333.8 g  $I_2O_5$  1 mol  $I_2O_5$  28.0 g  $CO$  1 mol  $CO$

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web Jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage  
text percent yield =  $\frac{\text{actual yield (g)}}{\text{theoretical yield (g)}} \times 100$  label 3 7 3

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