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web this data shows that there was a gain the mass of the lighter which shows that there was a certain amount of butane 0 38 grams 0 02 grams gas that was successfully collected during the gas collection there is a mild influence of barometric pressure during both trials as shown above in the raw data

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web 11 use avogadro's law to determine the number of moles of butane gas assume that butane is an ideal gas and that one mole has a volume of 22 4 l at stp $v_1 n_1 = v_2 n_2$ $0 0896 \text{ l} = 22 4 \text{ l} \cdot n$ $0 0040 \text{ moles}$ 12 experimental molar mass g/mole

of butane 0.23 g / 0.0040 moles = 58 g / mole. The molecular formula of butane is C_4H_{10} .

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web nov 3 2020 in this experiment the moles and mass will be required to determine the molar mass of butane in a butane lighter. Experimental: to begin the experiment the mass of a butane lighter was measured before using the butane gas to deplete the water to the 80ml mark on a graduated cylinder.

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web possible answer butane is extremely flammable and will ignite possible answer because the butane will not dissolve in the water it is possible for butane gas possible to form under water and get trapped possible answers 0.273 15 k and 1 atm 101.3 kPa C_4H_{10} 12.01 g / mol 4.48 04 g / mol

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web the pressure of the gas the pressure of the gas was calculated using Dalton's law of partial pressures by subtracting the pressure of the water vapor from the total pressure which in this case was the atmospheric pressure 760mm Hg the pressure of the water vapor was found by using the temperature of the water.

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