

REACTION KINETICS

MAKING CHEMICAL REACTIONS HAPPEN FASTER

There are a number of different things that we can change to make a chemical reaction faster. Here, we explain the concept of collision theory, and how it can be used to explain the effects of five different factors on the rate of a chemical reaction.

COLLISION THEORY



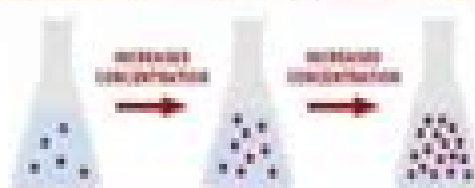
Correct orientation with enough energy

Wrong orientation

Wrong amount of energy

Collision theory states that, for a reaction to occur, particles must collide with the correct orientation and with sufficient energy for a reaction to occur. Different factors affect the rate of the reaction by affecting the frequency of particle collisions, and/or the proportion of collisions that have enough energy to react.

INCREASE CONCENTRATION OF REACTANTS



↑ FREQUENCY OF COLLISIONS

↑ % SUCCESSFUL COLLISIONS

Increasing the concentration of reactants in solution increases the rate of reaction as there are a greater number of particles available to react. This increases the frequency of collisions between particles.

INCREASE TEMPERATURE OF REACTION



↑ FREQUENCY OF COLLISIONS

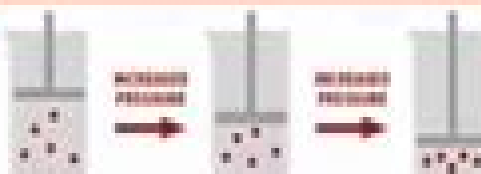
↑ % SUCCESSFUL COLLISIONS

Increasing the temperature increases the kinetic energy of particles. This increases the frequency of particle collisions, and a greater proportion of collisions will have the energy required to react.

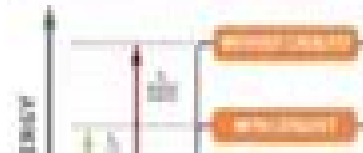
INCREASE SURFACE AREA OF REACTANTS



INCREASE PRESSURE OF REACTION



USE A CATALYST IN THE REACTION



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