

PIL TUTORIAL II

QUESTION:

The Chief Justice of Serenia has been chosen to sit on the bench of the ICJ as an ad hoc judge (a **judge** appointed by a special procedure for a specific project, case, or period only) in a case now before the Court. He has to deal with the mass of diverse materials submitted as part of the written pleadings.

Discussion.

1. Identify the category of the material.
2. Intro. of the category.
3. Element of the said category.
4. Relate back with the material.

(e) a treaty between the **two parties to the dispute;**

1. Identify the category of the material - primary source; material resource - treaty
2. Introduction of the category - what is treaty?

(FATHIN)

According to Article 2 of VCOLT: treaty means an international agreement concluded between States in written form and governed by international law, whether embodied in a single instrument or two or more related instruments and whatever its particular.

Oppenheim: international treaties are agreements of a contractual character, between states or organisations of states, creating a legal right and obligation between the parties. The efforts of the International Law Commission (ILC) and the United Nations have produced a number of significant multilateral treaties, such as the Vienna Convention on Diplomatic Relations 1961 and the Law of the Sea Convention 1982.

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N Noddings



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Table of Contents Questions Intl Law

1. Understanding the eBook Questions Intl Law
 - The Rise of Digital Reading Questions Intl Law
 - Advantages of eBooks Over Traditional Books
2. Identifying Questions Intl Law
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Questions Intl Law
 - User-Friendly Interface
4. Exploring eBook Recommendations from Questions Intl Law
 - Personalized Recommendations
 - Questions Intl Law User Reviews and Ratings

- Questions Intl Law and Bestseller Lists
- 5. Accessing Questions Intl Law Free and Paid eBooks
 - Questions Intl Law Public Domain eBooks
 - Questions Intl Law eBook Subscription Services
 - Questions Intl Law Budget-Friendly Options
- 6. Navigating Questions Intl Law eBook Formats
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 - Questions Intl Law Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Questions Intl Law
 - Highlighting and Note-Taking Questions Intl Law
 - Interactive Elements Questions Intl Law
- 8. Staying Engaged with Questions Intl Law
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Questions Intl Law
- 9. Balancing eBooks and Physical Books Questions Intl Law
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Questions Intl Law
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Questions Intl Law
 - Setting Reading Goals Questions Intl Law
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Questions Intl Law
 - Fact-Checking eBook Content of Questions Intl Law
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a byproduct of the process as iron ions the ion meta dissolves

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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and percent yield of this reaction will be calculated

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate CuSO_4 and 2 0153 grams of iron powder Fe the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper ii sulfate this reaction produces

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away cooper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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