

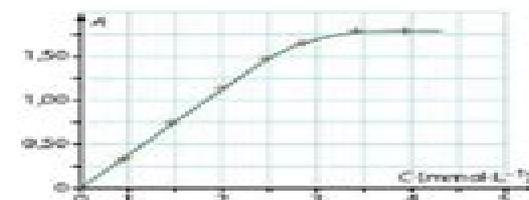
## Compétences travaillées

	Eco
<b>Absorbance ; loi de Beer-Lambert.</b>	<b>1</b>
Exploiter la loi de Beer-Lambert, la loi de Kohlrausch ou l'équation d'état du gaz parfait pour déterminer une concentration ou une quantité de matière. Citer les domaines de validité de ces relations.	<b>2</b>
<b>Conductance, conductivité ; loi de Kohlrausch.</b>	<b>3</b>
Exploiter la loi de Beer-Lambert, la loi de Kohlrausch ou l'équation d'état du gaz parfait pour déterminer une concentration ou une quantité de matière. Citer les domaines de validité de ces relations.	<b>4</b>
	<b>5</b>
<b>Spectroscopie UV-visible. Identification de groupes caractéristiques et d'espèces chimiques.</b>	<b>6</b>
Exploiter, à partir de données tabulées, un spectre UV-visible pour identifier un groupe caractéristique ou une espèce chimique.	<b>7</b>
	<b>8</b>
<b>Spectroscopie infrarouge. Identification de groupes caractéristiques et d'espèces chimiques.</b>	<b>9</b>
Exploiter, à partir de données tabulées, un spectre d'absorption infrarouge pour identifier un groupe caractéristique ou une espèce chimique.	<b>10</b>
	<b>11</b>

Liaison	Nombre d'ondes ( $\text{cm}^{-1}$ )	Intensité et commentaire
Liaison OH libre	Entre 3 590 et 3 650 $\text{cm}^{-1}$	Bande forte et fine
Liaison OH liée (liaison hydrogène)	Entre 3 200 et 3 600 $\text{cm}^{-1}$	Bande forte et large
Liaison C=O	Entre 1 625 et 1 820 $\text{cm}^{-1}$	Bande forte et de largeur moyenne
Liaison C=O des acides carboxyliques	Entre 1 660 et 1 740 $\text{cm}^{-1}$	Bande forte
Liaison C-H de CHO (carbone trigonal)	Entre 2 650 et 2 800 $\text{cm}^{-1}$	Bande moyenne
Liaison C-H (carbone tétraédrique)	Entre 2 800 et 3 000 $\text{cm}^{-1}$	Bande forte
Liaison C-O des alcools	Entre 1 000 et 1 150 $\text{cm}^{-1}$	Bande forte
Liaison OH des acides carboxyliques	Entre 2 500 et 3 300 $\text{cm}^{-1}$	Bande forte et large
Liaison C-O des acides carboxyliques	Entre 1 200 et 1 320 $\text{cm}^{-1}$	Bande forte

**Exercice 1 :** La courbe ci-dessous représente l'absorbance  $A$  en fonction de la concentration  $C$  en diode d'une gamme de solutions étalons. Dans les mêmes conditions de mesure que celles de la gamme étalon, une solution  $S$  de diode a une absorbance  $A_S = 1,25$ .

1. Énoncer la loi de Beer-Lambert.
2. Dans quel domaine de concentration le graphique traduit-il la loi de Beer-Lambert ? Justifier
3. Déterminer la concentration  $C$  en diode de la solution  $S$ .



**Exercice 2 :** Certains chewing-gums contiennent le colorant naturel E120 acide carminique. On introduit un chewing-gum dans de l'eau jusqu'à dissolution complète du colorant E120. On ajuste le volume de cette solution, notée  $S$ , à  $V_S = 50,0$  mL. Les absorbances  $A$  de solutions de différentes concentrations  $C$  en colorant E120 sont reportées dans le tableau ci-dessous :

$C$ ( $\text{mol.L}^{-1}$ )	$6,7 \times 10^{-6}$	$1,3 \times 10^{-5}$	$2,0 \times 10^{-5}$	$2,6 \times 10^{-5}$	$3,4 \times 10^{-5}$	$6,7 \times 10^{-5}$
$A$	0,06	0,12	0,17	0,25	0,30	0,59

**Données :**  $DJA(E120) = 2,5$  mg / kg de masse corporelle / jour et  $M(E120) = 492$  g.mol $^{-1}$ .

1. Déterminer la concentration  $C_S$  en colorant E120 de la solution  $S$  ( $A_S = 0,21$ ).
2. Calculer le nombre de chewing-gums qu'un élève de 60 kg devrait consommer pour atteindre la DJA du E120.

# Methodes Projectives Les Ancienne Adition

**Xiang Xie**



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