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## Chapter 5 Phase Transformations and Phase Diagrams Homework Solutions

### Concept Questions

1. If a transformation from solid to liquid occurs such that the solid and liquid are always in equilibrium, then the transformation is reversible.
2. A transformation occurs in a material, and the Gibbs free energy change in the material is negative. The transformation is spontaneous.
3. When  $\alpha$ -phase iron is in equilibrium with iron carbide, the iron atoms in the two phases have equal chemical potential.
4. Nickel and silver have lattice parameters of 0.352 nm and 0.408 nm, respectively; an electronegativity of 1.8 for each, an FCC structure, and a similar valence. Should they be able to form a continuous-solid solution? No
5. In a phase diagram, the liquidus line gives the temperature where the first solid forms upon cooling an alloy of given composition  $C_0$  from the liquid phase.
6. In a phase diagram, the solidus line gives the temperature where the first liquid forms upon heating an alloy of given composition  $C_0$  from the solid phase.
7. The lattice parameter of nickel is 0.352 nm, and for copper it is 0.363 nm. The electronegativities are 1.8 and 1.9, respectively, and they are both FCC metals. It is expected that there is continuous solid solubility of nickel in solid copper at elevated temperature.
8. When a liquid copper-nickel alloy is rapidly cooled, the center of the grains has a higher concentration of nickel atoms than the average concentration. This is coreing.

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**EW Minium**



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