

THEORIES OF COORDINATION COMPOUNDS

Valence bond theory (VB Theory)

Valence bond theory, primarily the work of Linus Pauling regarded bonding as characterized by the overlap of atomic or hybrid orbitals of individual atoms.

The postulates of valence bond theory

1. The central metal atom/ion makes available a number of vacant orbitals equal to its coordination number.
2. These vacant orbitals form covalent bonds with the ligand orbitals.
3. A covalent bond is formed by the overlap of a vacant metal orbital and filled ligand orbitals. This complete overlap leads to the formation of a metal ligand, σ (sigma) bond.
4. A strong covalent bond is formed only when the orbitals overlap to the maximum extent. This maximum overlapping is possible only when the metal vacant orbitals undergo a process called 'hybridisation'. A hybridised orbital has a better directional characteristics than an unhybridised one.

The following table gives the coordination number, orbital hybridisation and spatial geometry of the more important geometries.

CN	Geometry		Hybridization	Example
2	Linear		sp	$[\text{Ag}(\text{NH}_3)_2]^+$
4	Tetrahedral		sp^3	$[\text{Cd}(\text{NH}_3)_4]^{2+}$
4	square planar		sp^2d	$[\text{Cu}(\text{OH}_2)_4]^{2+}$
5	trigonal bipyramid		sp^3d	$\text{Fe}(\text{CO})_5$
5	Square pyramidal		sp^2d^2	$[\text{Mn}(\text{Cl})_5]^{2-}$

Mathematics Coordination

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Table of Contents Mathematics Coordination

1. Understanding the eBook Mathematics Coordination
 - The Rise of Digital Reading Mathematics Coordination
 - Advantages of eBooks Over Traditional Books
2. Identifying Mathematics Coordination
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mathematics Coordination
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mathematics Coordination
 - Personalized Recommendations
 - Mathematics Coordination User Reviews and Ratings
 - Mathematics Coordination and Bestseller Lists
5. Accessing Mathematics Coordination Free and Paid eBooks
 - Mathematics Coordination Public Domain eBooks
 - Mathematics Coordination eBook Subscription Services
 - Mathematics Coordination Budget-Friendly Options
6. Navigating Mathematics Coordination eBook Formats
 - ePub, PDF, MOBI, and More
 - Mathematics Coordination Compatibility with Devices
 - Mathematics Coordination Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mathematics Coordination
 - Highlighting and Note-Taking Mathematics Coordination
 - Interactive Elements Mathematics Coordination
8. Staying Engaged with Mathematics Coordination

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mathematics Coordination
9. Balancing eBooks and Physical Books Mathematics Coordination
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mathematics Coordination
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mathematics Coordination
- Setting Reading Goals Mathematics Coordination
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mathematics Coordination
- Fact-Checking eBook Content of Mathematics Coordination
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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