



# Reign Of Snakes

**Shasha Hu**



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$$4\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu} + 3\text{Ag} + 2\text{Cu}_2\text{AgNO}_3$$

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calculated mass of the product expected  $\times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  $4 \text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160 g

$4 \text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$  26 982 g  $4 \text{ mol Al} \times 1 \text{ mol Al}_2\text{O}_3$  c determine the percent yield

yield actual  $\times 100 = 260.0 \text{ g} \times 100 = 86.01$  theoretical 302.3 g

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weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12

c is the limiting reagent 4 33 g of h 2 are left over 13

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from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s

remains when the reaction is complete

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web the percent yield is calculated as follows  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent

yield theoretical  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  based on this definition we would expect a percent yield

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how much product will be produced from the limiting reactant calculate how much reactant s remains when the reaction is

complete

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$\text{CH}_3\text{OH}$  can be produced by the following reaction  $2 \text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$  calculate the theoretical yield of  $\text{CH}_3\text{OH}$  if 68.5 g of co is

reacted with 8.6 g of  $\text{H}_2$

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which reactant is limiting assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate what is the

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