

THERMODYNAMICS PROBLEMS

- What will be the sign on the value of the heat transferred in the case of an:
 - exothermic reaction.
 - endothermic reaction.
- What can happen to a system in which an internal energy change is occurring?
- A gas absorbs 28.5 J of heat and then performs 15.2 J of work. What is the change in internal energy of the gas?
- What are the signs of q and w for the following exothermic process at 1 atmosphere pressure and 370 Kelvin?

$$\text{H}_2\text{O}(g) \rightarrow \text{H}_2\text{O}(l)$$
- A system does 130. J of work on its surroundings while 70. J of heat are added to the system. What is the internal energy change for the system?
- What will be the volume change if 125 J of work is done on a system containing an ideal gas? The surroundings exert a constant pressure of 5.2 atmospheres.
- For the following reactions, performed at constant external pressure, is work done on the system by the surroundings, by the surroundings on the system, or is the amount of work negligible?
 - $\text{Sn}(s) + 2\text{F}_2(g) \rightarrow \text{SnF}_4(s)$
 - $\text{AgNO}_3(aq) + \text{NaCl}(aq) \rightarrow \text{AgCl}(s) + \text{NaNO}_3(aq)$
 - $\text{C}(s) + \text{O}_2(g) \rightarrow \text{CO}_2(g)$
 - $\text{SiH}_4(g) + \text{heat} \rightarrow \text{Si}(s) + 2\text{H}_2(g)$

- With the addition of heat and manganese dioxide as a catalyst, potassium chlorate at constant pressure will decompose according to the following equation:



Determine whether each of the values ΔE , ΔH , q , and w will be positive, negative, or unable to determine.

- Compare the quantities of heat released per mole of iron formed when oxides of Fe_2O_3 and Fe_3O_4 react with aluminum.



- The combustion of acetylene in a welding torch liberates intense heat. The equation for this reaction is:



How much heat, in joules, is liberated when 0.260 kg of C_2H_2 undergoes complete combustion?

- The standard enthalpy of formation for $\text{Cu}(\text{NO}_3)_2(s)$ is -302.6 kJ/mol . Write the thermochemical equation for the enthalpy of formation of this compound.
- Given the following enthalpy of formation equations:



- Sketch an enthalpy diagram for the formation of germanium dioxide in a one-step reaction.

Problems On Thermodynamics 4ed

Didier Musso



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